**Percent Error Notes Name \_\_\_\_\_\_\_\_\_\_\_**

=

When you are finding

the **percent error**

 you can use the

proportion:

**Predicted is what \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ happen**

**Actual is what \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ happened**

1. Chase worked in a lab helping pack material that was labeled to weigh 15 ounces. When one of his boxes was pulled and weighed it was 14.5 ounces. What was the percent error to the nearest tenth?

Actual =

Predicted =

2. Sally weighed an object on her balance and recorded a mass of 22.3 grams. The label on the object said that it should weigh 23.2 grams. What is the percent error to the nearest hundredth of a percent?

Predicted =

Actual =

3. Joshua uses his thermometer and finds the boiling point of ethyl alcohol to be 75o C.  He looks in a reference book and finds that the actual boiling point of ethyl alcohol is 80oC.  What is his percent error?

Predicted =

Actual =

**![C:\Documents and Settings\jainslie\Local Settings\Temporary Internet Files\Content.IE5\6W2FJPU3\MC900432687[1].png]()**

**Pause the video and try the problems on the back on your own!**

**Then press play and check your answers with a color pen.**

4. The literature value of the atomic mass of an isotope of nickel is 57.9 g/mol. If a laboratory experimenter determined the mass to be 59.6 g/mol, what is the percent error?

Predicted =

Actual =

5. The mass of one mole of oxygen gas is determined in an experiment to be 31.4 g/mol. Calculate the percent error, given that the literature value for this mass is 32.0 g/mol.

Actual =

Predicted =

6. A student calculates the density of a rock to be 2.2 g/ml and the accepted value is 2.7 g/ml. What is the student’s percent error?

Predicted =

Actual =